Worksheet #5

Name:

Period:

Seat#:

Directions: In each of the following equations, indicate the element that has been oxidized and the one that has been reduced. You should also label the oxidation state of each before and after the process. Also identify the oxidizing and reducing agent.

1)	2 Na + FeCl ₂ \rightarrow 2 NaCl + Fe
2)	$2 C_2 H_2 + 5 O_2 \rightarrow 4 CO_2 + 2 H_2 O$
3)	$2 \text{ PbS} + 3 \text{ O}_2 \rightarrow 2 \text{ SO}_2 + 2 \text{ PbO}$
4)	$2 H_2 + O_2 \rightarrow 2 H_2O$
5)	$Cu + HNO_3 \rightarrow CuNO_3 + H_2$
6)	$AgNO_3 + Cu \rightarrow CuNO_3 + Ag$